LP Monday, October 1st and Tuesday, October 2nd, 2018 CHAP 5

CALIFORNIA CONTENT STANDARDS: Science, Chemsitry

5.1.1 Identify the inadequacies in the Rutherford atomic model
5.1.2 Identify the new proposal in the Bohr model of the atom
5.1.3 Describe the energies and positions of electrons according to the quantum mechanical model
5.1.4 Describe how the shapes of orbitals, related to different sublevels differ.

Voc: **Copy Vocabulary # 5**
Define Words using text

Demo #6: **Hydrogen Balloon**
Exothermic Reaction: gas, heat, light

Chemistry Standard:
1e:students know how to relate the position of an element in the periodic table to its quantum electron configuaration
1i: students know the experimental basis for the development of the quantum theory or atomic structure and the historical importanceof the Bohr model
1j:students know that spectral lines are the result of transitions of electrons between energy levels

Lecture:
PPT: Electron Orbitals (Page 31 in NB)
Chapter 5.1 Electron Orbitals
Electron Orbitals Overhead
Copy down chart in notes

Class Work: **Electron Orbital Practice (1-4)**

WB: **Section 5.1, Page 43-44**

Home Work: **Page 132, 2 and 5**
Page 149, 22-29