Name _____ Date _____ Class _____

14 THE BEHAVIOR OF GASES

Vocabulary Review

Match the correct vocabulary term to each numbered statement. Write the letter of the correct term on the line.

	Column A	Column B
_____ **1.**	At constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures of the component gases.	**a.** combined gas law
_____ **2.**	The volume of a fixed mass of gas is directly proportional to its Kelvin temperature if the pressure is kept constant.	**b.** ideal gas constant (R)
_____ **3.**	The rate of effusion of a gas is inversely proportional to the square root of its molar mass.	**c.** diffusion
_____ **4.**	the contribution each gas in a mixture makes to the total pressure of that mixture	**d.** compressibility
_____ **5.**	a measure of how much the volume of matter decreases under pressure	**e.** Boyle's law
_____ **6.**	For a given mass of gas at constant temperature, the volume of the gas varies inversely with pressure.	**f.** partial pressure
_____ **7.**	the tendency of molecules to move toward areas of lower concentration until the concentration is uniform throughout	**g.** Dalton's law of partial pressures
_____ **8.**	$\frac{P_1 \times V_1}{T_1} = \frac{P_2 \times V_2}{T_2}$	**h.** effusion
_____ **9.**	the escape of a gas through a tiny hole in a container of gas	**i.** Charles's law
_____ **10.**	8.31 (L•kPa)/(K•mol)	**j.** Graham's law of effusion