11	2.1 H																
•	1.0	1.5	7										3A 2.0	4A 2.5	5A	6A	7
	Li	Be	}										B B	C	3.0 N	3.5 O	4
	0.9 Na	1.2 Mg											1.5 Al	1.8 Si	2.1 P	2.5 S	3. C
, ,	0.8 K	1.0 Ca	1.3 Sc	1.5 Ti	1.6 V	1.6 Cr	1.5 Mn	1.8 Fe	1.8 Co	1.8 Ni	1.9 Cu	1.6 Zn	1.6 Ga	1.8 Ge	2.0 As	2.4 Se	2. B
	0.8 Rb	1.0 Sr	1.2 Y	1.4 Zr	1.6 Nb	1.8 Mo	1.9 Tc	2.2 Ru	2.2 Rh	2.2 Pd	1.9 Ag	1.7 Cd	1.7 In	1.8 Sn	1.9 Sb	2.1 Te	2.
	0.7 Cs	0.9 Ba	1.1-1.2 La-Lu	1.3 Hf	1.5 Ta	1.7 W	1.9 Re	2.2 Os	2.2 Ir	2.2 Pt	2.4 Au	1.9 Hg	1.8 T1	1.8 Pb	1.9 Bi	2.0 Po	2. A
4 1	0.7 Fr	0.9 Ra	1.1-1.7 AcLr					-									_
	0.7 Cs 0.7	0.9 Ba 0.9	Y 1.1-1.2 La-Lu 1.1-1.7	Zr	Nb	Mo 1.7	Tc 1.9	Ru 2.2	Rh 2.2	Pd 2.2	Ag 2.4	Cd	In 1.8	Sn 1.8	Sb	Te 2.0	

1.	Why is no electronegativity shown for element 10?

2. Calculate the electronegativity difference for the atoms that are bonded in the following diatomic molecules. Then tell whether the bond is nonpolar covalent, polar covalent, or ionic and which atom has the greater share of the bonding electrons.

Formula	Electronegativity Difference	Type of Bond	Atom With Greater Electron Share
NO			
MgO			
Br ₂			
LiH			

Electron Dot Structures

Molecular bonding is represented by electron dot structures in which valence electrons are shown as dots around the chemical symbols of the atoms. Draw electron dot structures for each of the following compounds or ions.

CIF

Cl₂O

 C_2Cl_2

 NH_3

 NF_4 +

 C_2Cl_4

 CO_2

 O_3