LP Friday, May 10th and Monday, May 13rd , 2018 Equilibrium Chapter 18

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Chemistry Content Standards:

Chemical equilibrium is a dynamic process at the molecular level.

Standard 9a: Students know how to use LerChatelier's principle to predict the effect of changes in concentration, temperature, and pressure.

Standard 9b: Students know equilibrium is established when forward and reverse reaction rates are equal.

Standard 9c: Students know how to write and calculate an equilibrium constant expression for a reaction.

18.1.1: Describe how to express the rate of a chemical reaction
18.1.2: Identify four factors that influence the rate of a chemical reaction

Stamp: **All unstamped work**

Lecture: **Chapter 18.1, Rates of Reactions**

Chemical Kinetics PPT

- The equilibrium constant, K_c
- Gas equilibrium constant, K_p
- Relationship between K_c and K_p
- Concentration changes
- The reaction Quotient

Chapter 18.2, Reversible Reactions and Equilibrium

Equilibrium and Le Chatelier's Principle PPT

- Equilibrium constant, K_{eq}
- Reversible reactions
- Le Chatelier's Principle

Notes: **Chemical Kinetics** page 63 NB

Class work: Students should:

1. Equilibrium Practice Problems
2. Practice Problems 18.1
3. Practice Problems 18.2

Homework: None