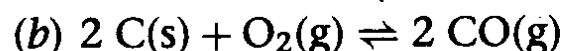
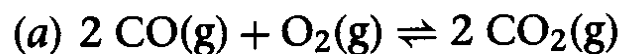
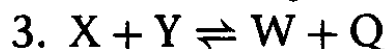
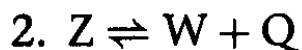
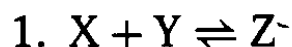


EXAMPLE 1 Determine the equilibrium constant expression for each of the following reactions:



EXAMPLE 2 (a) Write equilibrium constant expressions for equations 1 to 3 below. (b) Determine the relationship of the K value for equation 3 to those of equations 1 and 2.



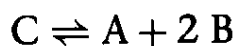
EXAMPLE 3 Calculate the value of the equilibrium constant for the reaction



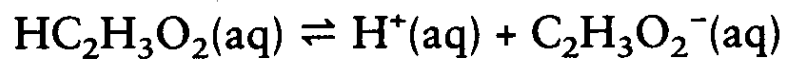
if the concentrations at equilibrium are $[\text{A}] = 2.0 \text{ M}$, $[\text{B}] = 1.5 \text{ M}$, and $[\text{C}] = 0.010 \text{ M}$.

EXAMPLE 4

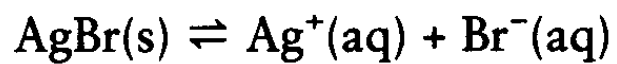
(a) Using the data of Example 3, calculate the value of the equilibrium constant for the reaction



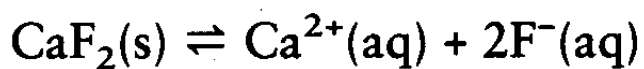
(b) What is the relationship between the value of K in Example 3 and the value of this K ?



$$K_{eq} = \underline{\hspace{4cm}}$$



$$K_{sp} = \underline{\hspace{4cm}}$$



$$K_{sp} = \underline{\hspace{4cm}}$$

