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Market Equilibrium



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Chemistry Content Standards:

Chemical equilibrium is a dynamic process at the molecular level.

Standard 9a: Students know how to use LeChatelier's principle to predict the effect of changes in concentration, temperature, and pressure.

Standard 9b: Students know equilibrium is established when forward and reverse reaction rates are equal.

Standard 9c: Students know how to write and calculate an equilibrium constant expression for a reaction.

- 18.2.1: Describe how the amounts of reactants and products change in a chemical system at equilibrium
12.2.2: Identify three stresses that can change the equilibrium position of a chemical system
18.2.3: Explain what the value of K_{eq} indicates about the position of equilibrium.

Stamp Homework:

Equilibrium Practice Problems 18.1 and 18.2

Lab

Lab # 16: Does Steel Burn? Page 64 in NB
Purpose: To determine whether steel will burn.
Student hand out for lab write up

Class Work: Complete **Lab # 16** and lab write up. Answer all questions to achieve full credit

Complete **Problem Solving Chapter 18:** Reaction Equilibrium

Home Work: Prepare Notebooks for last check next class